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Chemical Reactions. Using
the following equation: $\text{Fe}_2\text{O}_3 (\text{s}) + 3\text{H}_2 (\text{g}) \rightarrow 2\text{Fe} (\text{s})$

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+ 3H₂O(l) How many moles of Fe can be made from 6 moles H₂? (This is a one step conversion using mole ratio)

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Stoichiometry Stoichiometry
is the practice of
predicting the amount of
product or reactant in a
chemical equation based on a
known amount of one of the
other products or reactants.

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The ability to do these calculations is the culmination of all of the basic skills learned throughout the first semester.

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Stoichiometry: Mixed

Problems (KEY) 1) $N_2 + 3H_2 \rightarrow 2NH_3$?

What volume of NH_3 at

STP is produced if 25.0 of

N_2 is reacted with an excess

of H_2 ? 3 3 3 2 3 2 2 2 40.0L

NH_3 1mol NH_3 22.4L NH_3 1mol N

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2mol NH 28.0g N 25.0g N 1mol
N x x x = 2) 2KClO₃ ? 2KCl +
3O₂ If 5.0g of KClO₃ is
decomposed, what volume of
O₂ is produced at STP? 2

~~Stoichiometry: Mixed
Problems (KEY)~~

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Mixed Stoichiometry Problems

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. a).

How many moles of H_2 would be required to produce 5.0 moles of water?

5.0 moles water. b). What mass of H_2O is formed when H_2 reacts with 384 g of O_2 ? 432g H_2 .

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2. $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$. a). Balance this equation. Look above. b).

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Learn how to use mole ratios derived from balanced chemical equations to

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calculate amounts of
substances consumed and
produced in chemical
reactions.

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calculation of reactants and

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products in chemical reactions.. Stoichiometry is founded on the law of conservation of mass where the total mass of the Page 11/21

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