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atoms that constitute a particular chemical compound, using element symbols and numbers. 397 People Used

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Characteristics of both bonds: * Occur between 2 atoms * Composed of 2 electrons * Have both ionic and covalent characteristics * Together = 100% * Both bonds are measured on an electronegativity scale * Both contain a nonmetal * Chemical bonds * Are

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determined by using the “ magic number ” (1.67) * Have bond angle and bond axis

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Metals (cations) Nonmetals(anions)
Cations An atoms that has lost or gained electrons creating a charge. Loses

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electrons Anions Gains electrons Charge
that ions normally make when bonding
Most atoms need 8 valence electrons in
their outermost shells in order to be stable.
To show the charge of multivalent ions.

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Chemical Bonding Study Guide A. Fill in the Blank 1. The electrons involved in the formation of a chemical bond are called 2. A chemical bond that results from the electrostatic attraction between positive and negative ions is called a(n) 3. If electrons involved in bonding spend most of the time closer to one atom rather than

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1. for most combinations of atoms a chemical bond will occur if the potential energy is reduced by the change a. most atoms have lower potential energy in bonds 2. chemical bonding is often an energy releasing process 3. breaking of

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chemical bonds is often an energy absorbing process

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links the two atoms together into something called a molecule. Making a...

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What is a chemical bond? a mutual electrical attraction between the nuclei and valence e- of different atoms that binds the atoms together What happens to

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the valence electrons in an ionic bond vs. a covalent bond? In an ionic bond, they are transferred.

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the lesson 2 you studied about the classification of elements and the variation in atomic properties. You know that molecules are obtained by the combination of two or more than two atoms of the same or different elements. In this lesson you will study

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Covalent bonding represents a situation of about equal sharing of the electrons between nuclei in the bond. Covalent bonds are formed between atoms of approximately equal electronegativity. Because each atom has near equal pull for the electrons in the bond, the electrons are not completely transferred from one atom

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to another.

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November 2016 A molecule with covalent
bond o Formula unit with ionic bond *

Molecules: formed by covalent bonds *

Lattice energy is the energy released in the
formation of an ionic compound.

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